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Solutions of amount of substance per unit volume of solution. In chemistry, the most commonly used unit for molarity is the number of moles per litre, having the unit symbol mol/L or mol·dm⁻³ in SI unit.

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Chemistry Dimensions 1

A liter is a little larger than 1 US quart in volume. (Table 1.4) gives approximate equivalents for some of the units used in

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chemistry.) As shown in Figure 1.9 “The Liter”, a liter is also 1,000 cm³. By definition, there are 1,000 mL in 1 L, so 1 milliliter and 1 cubic centimeter represent the same volume. 1 mL = 1 cm³

Chapter 1: Measurements in Chemistry - Chemistry

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Solution, in chemistry, a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is commonly applied to the liquid state of matter, but solutions of gases and solids are possible.

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solution | Definition & Examples | Britannica

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ensure that the particle has zero probability of being at the walls or outside the box.

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Molar concentration (also called molarity, amount concentration or substance concentration) is a measure of the concentration of a chemical species, in particular of a solute in a

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solution, in terms of amount of substance per unit volume of solution. In chemistry, the most commonly used unit for molarity is the number of moles per litre, having the unit symbol mol/L or mol·dm⁻³ in SI unit.

Molar concentration - Wikipedia

Solution: (i) Linear momentum can be

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written as “mv”. Dimensions of Linear momentum: $mv = [MLT^{-1}]$ (ii)
Frequency can be written as “ $1/T$ ”, where T is time. Dimensions of Frequency: $1/T = [M^0 L^0 T^{-1}]$ (iii) All units of pressure represent some ratio of force to area. So, dimensions of pressure: Force/Area =.

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The magnitudes of physical quantities can be added together or subtracted from one another only if they have the same dimensions. So, the dimensions of d will be $[T]$ Now Velocity has the dimension $[LT^{-1}]$ $[L T^{-1}]$, Now each of the quantity which are being added

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on the RHS should also have same dimension

dimensional analysis practice problems

A dispersion is a system in which distributed particles of one material are dispersed in a continuous phase of another material. The two phases may

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be in the same or different states of matter.. Dispersions are classified in a number of different ways, including how large the particles are in relation to the particles of the continuous phase, whether or not precipitation occurs, and the ...

Dispersion (chemistry) - Wikipedia

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